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Stoichiometry - Limiting \u0026amp;#x27E6

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~~Excess Reactant, Theoretical
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Stoichiometry - Chemistry for
Massive Creatures: Crash Course
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Stoichiometry Problems - College
Chemistry How to Calculate
Percent Yield and Theoretical
Yield The Best Way - TUTOR
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Notes & Practice
Questions Part 1 PCAT General
Chemistry Review Test Prep
Study Guide Course
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Conversions Made Easy: How to
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Moles Stoichiometry Made Easy:
The Magic Number Method
Limiting Reactant Practice~~

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Problem (Advanced)

Limiting Reactant Practice

Problem Limiting Reagents and
Percent Yield Stoichiometry

Tutorial: Step by Step Video +
review problems explained |

Crash Chemistry Academy

~~Stoichiometry: Converting Grams
to Grams~~ *Easiest way to solve*

*limiting reagent problems - ABCs
of limiting reagent* ~~Hydrocarbon~~

~~Power! Crash Course Chemistry
#40~~ *Introduction to Oxidation*

Reduction (Redox) Reactions

MCAT Test Prep General

Chemistry Review Study

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Equations \u0026 Formulas -

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Problems HOW TO STUDY FOR

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stoichiometry • mole ratio
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CHEMISTRY I (TESC 141) STUDY
GUIDE MOLES/ STOICHIOMETRY

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Mole-a unit of measurement that expresses the amount of atoms, molecules or some other unit. The number of items in one mole is commonly referred to Avogadro's number which equals 6.022×10^{23} .

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Equations • Number of Atoms •
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Moles • Mass o Law of
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Monday (1/27): In-class: Naming,
Formula, and Balancing Review;
HW: Finish Review, Watch
Stoichiometry Lecture Video and
Complete Guided Notes (p 1-3)
Due Tomorrow Tuesday (1/28): In-
class: Note Check, Funsheets A-E;
HW: Study for Quiz Friday
Wednesday (1/29): In-class:

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Answers A-E; HW: Watch
Percent Yield and Limiting
Reactant Lecture Video and
Complete Guided Notes (p 4-6)
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Stoichiometry Section 11.1 What
is stoichiometry? 1. true 2. true 3.
false 4. true 5. true 6. 2, 2, 64.10

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7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4,
4, 72.08 10. methanol and
oxygen gas 11. carbon dioxide
and water 12. 160.10 g 13.
160.10 g 14. They are equal. 15.
A mole ratio is a ratio between ...

VIBRATIONS AND WAVES

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Pages. Stoichiometry Section 11.1

What is stoichiometry? In your
textbook, read about

stoichiometry and the balanced
equation. For each statement
below, write true or false.

_____ true _____ 1. The
study of the quantitative
relationships between the
amounts of reactants used and
the amounts of products formed
by a chemical reaction is called

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stoichiometry.

_____true_____ 2.

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