

Heterolytic Fragmentation Of Organic Molecules 1st Edition

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Heterolytic fragmentation is a widespread but neglected class of organic reactions. It involves the regulated cleavage of molecules containing certain combinations of atoms such as carbon, oxygen, nitrogen, sulfur, phosphorus, silicon, boron and halogens. Fragmentation reactions are useful in degradation and structure elucidation, some are also of preparative value.

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The study of heterolytic fragmentations contributes heavily to the understanding of heteroelectronic effects in organic chemistry. In addition, sometimes the most expedient way to incorporate (in a synthesis) a particular framework or array of functional groups uses a fragmentation pathway, or a fragmentation may be a mandatory step in determining a structure.

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Homolytic vs. Heterolytic Fragmentation. Most organic transformations involve the movement of electron pairs (heterolytic reactions). There are a few important addition reactions, however, in which the electron reconfiguration involves the movement of single electrons. Whereas heterolytic bond cleavage leads to ion pairs, homolytic bond cleavage results in unpaired electrons \cdot or free radicals.

[Homolytic vs. Heterolytic Fragmentation](#)

In chemistry, heterolysis or heterolytic fission (from Greek $\eta\tau\epsilon\rho\omicron\varsigma$, heteros, "different", and $\lambda\upsilon\sigma\iota\varsigma$, lysis, "loosening") is the process of cleaving a covalent bond where one previously bonded species takes both original bonding electrons from the other species. During heterolytic bond cleavage of a neutral molecule, a cation and an anion will be generated.

[Heterolysis \(chemistry\) - Wikipedia](#)

Fragmentation is a type of chemical dissociation, in which the removal of the electron from the molecule results in ionization. Removal of electrons from either sigma bond, pi bond or nonbonding orbitals causes the ionization. This can take place by a process of homolytic cleavage/ homolysis or heterolytic cleavage/ heterolysis of the bond. Relative bond energy and the ability to undergo favorable cyclic transition states affect the fragmentation process.

[Fragmentation \(mass spectrometry\) - Wikipedia](#)

The origin of fragmentation patterns The formation of molecular ions When the vaporised organic sample passes into the ionisation chamber of a mass spectrometer, it is bombarded by a stream of electrons. These electrons have a high enough energy to knock an electron off an organic molecule to form a positive ion.

[FRAGMENTATION PATTERNS IN THE MASS SPECTRA OF ORGANIC ...](#)

Heterolytic fission, also known as heterolysis, is a type of bond fission in which a covalent bond between two chemical species is broken in an unequal manner, resulting in the bond pair of electrons being retained by one of the chemical species (while the other species does not retain any of the electrons from the bond pair).

[Homolytic & Heterolytic Fission of Covalent Bonds ...](#)

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Two types of heterolytic fragmentations that have been reported in the literature but have received little theoretical attention are (1) extended fragmentations – heterolytic fragmentations that involve a chain of more than five atoms – and (2) divergent fragmentations – heterolytic fragmentations that involve the formation of two (or more) distinct products from a single substrate. 4 Extended fragmentations are unsurprisingly rare due to the complexity of orchestrating many bond ...

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Four-membered heterocycles most often give fragments containing two ring atoms with their respective ligands. However, [3 + 1] fragmentation is well known, giving an atom (such as S) or stable species (such as SO₂) and often a three-membered ring. Examples are found in Schemes 4 and 5.

[Fragmentation Reaction - an overview | ScienceDirect Topics](#)

For most organic compounds the threshold energy for EI is about ~ 10 eV. After ionization of the molecule, the residual energy from the collision may cause the molecular ion to fragment into neutral pieces and smaller fragment ions. This fragmentation can be through either homolytic cleavage or heterolytic cleavage. The molecular ion is a radical cation, but the fragment ions may either be radical cations or carbocations, depending on the nature of the neutral fragment.

[thermodynamics - Heats of formation of neutral molecules ...](#)

A type of dissociation where the electrons are removed from the molecule to form ions is called fragmentation. This fragmentation can be carried out by heterolytic or homolytic cleavage of the bond. Thus, the bond energy and its function to undergo cyclic transition states cause a major effect on the fragmentation process.

[Learn About Fragmentation In Mass Spectrometry \(Ms ...](#)

Fragmentation is a type of chemical dissociation, in which removal of the electron from molecule result in ionization. Removal of electrons from either sigma bond, pi bond or nonbonding orbitals causes the ionization. [2] That can take place by a process of homolytic cleavage/ homolysis or heterolytic cleavage/ heterolysis of the bond. Relative bond energy and the ability to undergo favorable cyclic transition states affect the fragmentation process.

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