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**7.B Mole and Mass Avogadro's
Number, The Mole, Grams,
Atoms, Molar Mass**

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Calculations - Introduction

Converting Between Grams and

Moles Mole Ratio Practice

Problems **Concept of Mole -**

Part 1 | Atoms and Molecules

| Don't Memorise Empirical

Formula \u0026 Molecular

Formula Determination From

Percent Composition **Molality**

Practice Problems -

Molarity, Mass Percent, and

Density of Solution Examples

~~Introduction to Moles~~

Stoichiometry Basic

Introduction, Mole to Mole,

Grams to Grams, Mole Ratio

Practice Problems *Mole*

Conversions Made Easy: How

to Convert Between Grams and

Moles How To Convert Grams

To Moles - VERY EASY!

Counting Atoms: Intro to

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~~Key~~ Part 2 ~~Converting~~
~~Grams to Moles Using Molar~~
~~Mass | How to Pass Chemistry~~
~~The Mole: Avogadro's Number~~
~~and Stoichiometry~~
Introduction to Combustion
Analysis, Empirical Formula
\u0026 Molecular Formula
Problems Dalton's Law of
Partial Pressure Problems
\u0026 Examples - Chemistry
~~Stoichiometry Mole to Mole~~
~~Conversions - Molar Ratio~~
~~Practice Problems~~ Step by
Step Stoichiometry Practice
Problems | How to Pass
Chemistry Balancing Chemical
Equations Practice Problems
~~Concept of Mole | Avogadro's~~
~~Number | Atoms and Molecules~~
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CHEMISTRY WORKSHEET # 6

MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP.

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CHEMISTRY WORKSHEET # 2: THE

MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using

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Key the factor-label technique.

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Teacher Sep*

A mole of carbon contains
 6.02×10^{23} atoms of carbon.

A mole of helium contains
 6.02×10^{23} atoms of helium.

A mole of sodium contains
 6.02×10^{23} atoms of sodium.

A mole of gold contains 6.02
 $\times 10^{23}$ atoms of gold.

However, we have learned
about seven elements that
exist as diatomic

molecules— H_2 , N_2 , O_2 , F_2 ,
 Cl_2 , Br_2 , and I_2

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THE MOLE AS A UNIT OF

MASS Now that you know how to
find the mass of one mole of
a substance (molar mass) you
can easily find the mass
of several moles or the mass

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Key
of a fraction of a mole
using the factor-label
technique. 1 mol = a molar
mass of an atom/molecule
(g/mol) EXAMPLE: What is the
mass of 5.00 moles of water
H₂O? STEP 1: 2 H = 2 x (
1.01) = 2.02 STEP 2: ? grams
H₂O = 5.00 moles H₂O x
18.02 g H₂O = 90.10 ...

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6 CHEMISTRY WORKSHEET # 4
MOLAR MASSES AND AVOGADRO'S
NUMBER Now that you know two
definitions of a mole (a
molar mass and an Avogadro's
number of particles) you
can combine these two

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Key
definitions into one
problem. EXAMPLE: How many
molecules are there in 90.1
grams of water?
 $2 \text{ H} = 2$
 $\times (1.01) = 2.020 = 1 \times$
 $(16.00) = 16.00 \text{ H}_2\text{O} = 18.02$
g? molecules $\text{H}_2\text{O} = 90.1 \text{ g}$
 $\text{H}_2\text{O} \times 1 \text{ mole H}_2\text{O} \times 6.02$
 $\times 10^{23} \text{ molecules H}_2\text{O} =$
 $3.01 \times 10^{24} \text{ molecules H}_2$
...

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WORKSHEET # 6 MIXED MOLE
PROBLEMS (GRAMS, MOLECULES,
AND LITERS) You now know
three things a mole can be:
a molar mass, 6.02×10^{23}
molecules and, for a gas,
22.4 liters at STP. We can
use this

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CHEMISTRY WORKSHEET #5 MOLE
PROBLEMS-MOLAR VOLUME OF A
GAS We Have Learned That A
Mole Is A Mass Of Material

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Key
And Number Of Particles. A Mole Can Also Be A Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The Physical

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