

Chapter 11 Chemical Reactions Work Answers

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Chapter 11 Chemical Reactions Work
Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$ D

Chapter 11: Chemical Reactions

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Figure 11.1 Coral reefs, such as this one at the Palmyra Atoll National Wildlife Refuge, are vital to the ecosystem of earth's oceans. The health of coral reefs and all marine life depends on the specific chemical composition of the complex mixture known as seawater. (credit: modification of work by "USFWS – Pacific Region"/Wikimedia Commons)

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano.Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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Assume there are no side reactions or auxiliary reactions. From Eqs. 11.5.9 and 11.5.10, calculate the standard molar internal energy of combustion of n-hexane at $\{(298.15\text{K})\}$. (p) From Eq. 11.5.13, calculate the standard molar enthalpy of combustion of n-hexane at $\{(298.15\text{K})\}$. 11.8

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